

Knox J H. A new mechanism for the low temperature oxidation of hydrocarbons in the gas phase. *Combust. Flame* 9:297-310, 1965.
 [Department of Chemistry, University of Edinburgh, Scotland]

The gas phase oxidation of an alkane (C₂-C₅) produces the conjugate alkene as the major product in the temperature range 300-400°C. Involvement of this alkene is crucial to the subsequent progress of the reaction and enables initially formed HO₂ radicals to be converted to the much more reactive OH radicals. [The SCI[®] indicates that this paper has been cited in over 100 publications since 1965, making it the most-cited paper ever published in this journal.]

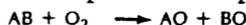
John H. Knox
 Department of Chemistry
 University of Edinburgh
 Edinburgh EH9 3JJ
 Scotland

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"The new mechanism for low temperature oxidation of alkanes proposed in this paper arose out of investigations which started under R.G.W. Norrish in 1947, when I went up to Cambridge to study for my PhD degree. We used a flow system to study the 'cool-flame oxidation of propane,' condensing out liquid products at -80°C and collecting uncondensed gaseous products in a bulb. The products were analysed by standard volumetric methods or by the then current gas handling techniques.¹ Both methods were laborious, time-consuming, insensitive, and relatively unselective. The objective of this work, to put it bluntly, was to substantiate the 'aldehyde theory' proposed by Norrish² and to discredit the rival 'peroxide theory.'³ Unfortunately, things did not work out quite as hoped! Neither theory showed much correlation with reality. Higher aldehydes, which Norrish proposed were the key autocatalytic agents, were minor products, and the 'peroxides,' favoured by the rival school, were mainly condensation products of H₂O₂ and aldehydes, probably formed in the cold trap.

"It was clear that no real progress could be made with the analytical methods then in existence. Something radically new was required, and gas chromatography (GC) provided the answer. On my appointment to a lectureship in Edinburgh, I applied GC to study the mechanism of the isothermal combustion of hydrocarbons between 300 and 400°C.⁴ A key discovery was soon made.⁵ At

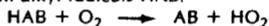
alkene conversions below about one percent, the major products (> 80 percent) from oxidation of ethane, propane, and isobutane were always the alkenes having the same number of C-atoms as the parent alkane. Oxygenated products, while formed in parallel, were always present in very small quantities relative to the alkene. If any product was responsible for the autocatalysis of alkane combustion it was most likely to be alkene. It was also shown that carbonyl compounds were the primary products of oxidation of the alkenes⁶ so the main steps in the combustion of an alkane, HABH, were:



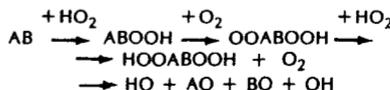
(AB = alkene; AO, BO = aldehyde or ketone)

"It was further established that the composition of the oxygenated products were strongly dependent on the nature of the walls of the reaction vessel⁶ and therefore that the elaborate attempts which were being made around 1965 to explain their formation by homogeneous unimolecular reactions of peroxy radicals (HABOO) were misconceived.

"The paper assembles these ideas. HO₂ radicals are first formed by hydrogen abstraction by oxygen from alkyl radicals HAB.



It is proposed that these then react mainly by addition to AB followed by further addition of oxygen, reaction with a second HO₂ radical, and final homogeneous decomposition of the diperoxide HOOABOOH to give the carbonyl compounds AO and BO. The main steps are then:



"In this way HO₂ radicals are converted to OH radicals which can abstract H atoms and propagate oxidation chains via the alkene as intermediate.

"The wide citation of this paper probably arises because it gives a clear and concise account of a mechanism of alkane combustion based upon detailed chemical analysis, which explained all the major features of these complex reactions as known in 1965. Furthermore, its conclusions are still broadly valid today."

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Proc. Roy. Soc. A 221:151-70, 1954.

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3. Ubbelohde A R. Investigations on the combustion of hydrocarbons. I—The influence of molecular structure on hydrocarbon combustion; and II—Absorption spectra and chemical properties of intermediates.

Proc. Roy. Soc. A 152:354-77; 378-402, 1935.

4. Falconer I W & Knox J H. The high-temperature oxidation of propane. *Proc. Roy. Soc. A* 250:493-513, 1959.

5. Knox J H. Some features of the oxidation of propane and ethane at 318°C. *Trans. Faraday Soc.* 55:1362-74, 1959.

6. Hay J, Knox J H & Turner J M C. Homogeneous and heterogeneous processes in the gas-phase oxidation of isobutane and isobutene. *Tenth Symposium (International) on Combustion at the University of Cambridge, Cambridge, England, August 17 to 21, 1964.* Pittsburgh, PA: Combustion Institute, 1965. p. 331-40.